

Modified Enlarged 24pt
OXFORD CAMBRIDGE AND RSA EXAMINATIONS

Monday 19 May 2025 – Morning

GCSE (9–1) Chemistry A (Gateway Science)
J248/01 (Foundation Tier)

Time allowed: 1 hour 45 minutes
plus your additional time allowance

YOU MUST HAVE:

a ruler (cm/mm)
the Data Sheet for GCSE (9–1)
Chemistry A (with this document)

YOU CAN USE:

a scientific or graphical calculator
an HB pencil
a model for Question 22(c)

Please write clearly in black ink.

Centre number

--	--	--	--	--

Candidate number

--	--	--	--

First name(s) _____

Last name _____

READ INSTRUCTIONS OVERLEAF



INSTRUCTIONS

Use black ink. You can use an HB pencil, but only for graphs and diagrams.

Write your answer to each question in the space provided. If you need extra space use the lined pages at the end of this booklet. The question numbers must be clearly shown.

Answer ALL the questions.

Where appropriate, your answer should be supported with working. Marks might be given for using a correct method, even if your answer is wrong.

INFORMATION

The total mark for this paper is 90.

The marks for each question are shown in brackets [].

Quality of extended response will be assessed in questions marked with an asterisk (*).

ADVICE

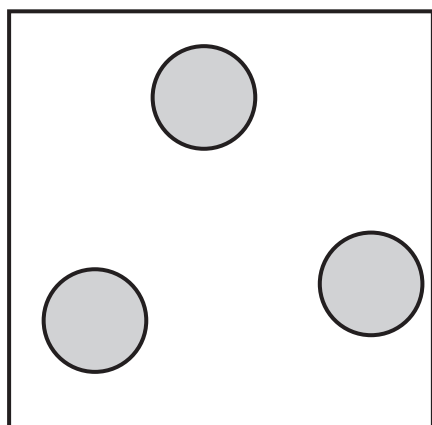
Read each question carefully before you start your answer.

SECTION A

You should spend a **MAXIMUM** of **30 MINUTES** plus your additional time allowance on this section.

Write your answer to each question in the box provided.

- 1 Which state of matter is shown in the particle model diagram? [1]



A Aqueous

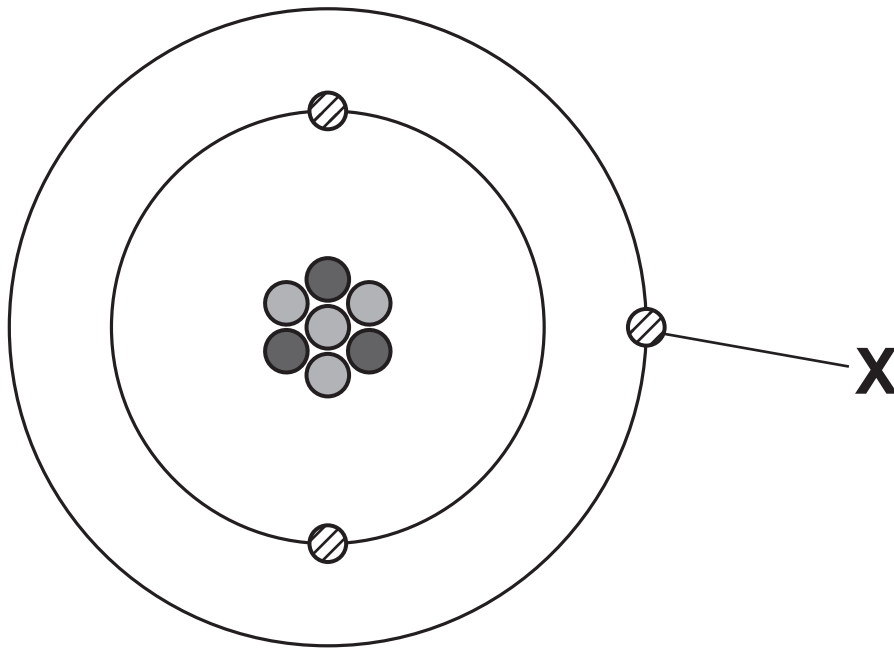
B Gas

C Liquid

D Solid

Your answer

2 What is the part of the atom labelled X? [1]



A Electron

B Neutron

C Nucleus

D Proton

Your answer

3 What is the typical RADIUS of an atom? [1]

A $1 \times 10^{-10} \text{ m}$

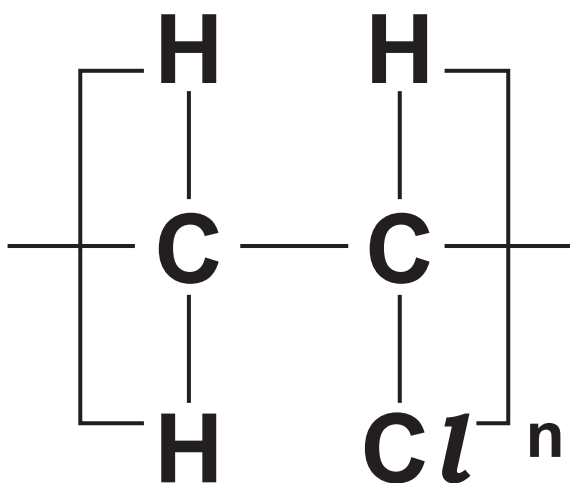
B $1 \times 10^{-1} \text{ m}$

C $1 \times 10^1 \text{ m}$

D $1 \times 10^{10} \text{ m}$

Your answer

4 Which type of molecule is shown in the diagram? [1]



A Alkane

B Alkene

C Monomer

D Polymer

Your answer

☐

5 What is meant by ACTIVATION ENERGY? [1]

- A The maximum amount of energy needed to end a reaction.**
- B The maximum amount of energy needed to start a reaction.**
- C The minimum amount of energy needed to end a reaction.**
- D The minimum amount of energy needed to start a reaction.**

Your answer

6 How did Mendeleev order the elements in his version of the Periodic Table? [1]

- A Decreasing atomic mass**
- B Decreasing atomic number**
- C Increasing atomic mass**
- D Increasing atomic number**

Your answer

- 7 Fluorine has a melting point of -220°C and a boiling point of -188°C .**

What is the state of fluorine at 25°C ? [1]

A Aqueous

B Gas

C Liquid

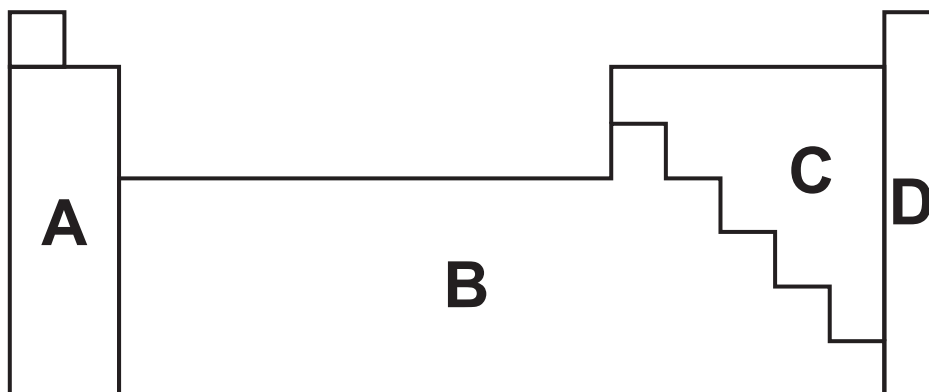
D Solid

Your answer

☐

8 An element has an electronic arrangement 2.8.1.

Which section of the Periodic Table is the element in? [1]



Your answer

9 Copper reacts with oxygen to make copper oxide.

Which mass of oxygen reacts with 127 kg of copper to make 159 kg of copper oxide? [1]

copper	+	oxygen	→	copper oxide
127 kg				159 kg

A 16 kg

B 32 kg

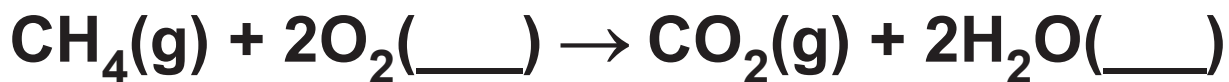
C 64 kg

D 128 kg

Your answer

10 Methane, CH₄, burns in oxygen, O₂.

Carbon dioxide, CO₂, and water vapour, H₂O, are made.



What are the missing state symbols in the equation? [1]

	O ₂	H ₂ O
A	g	g
B	g	l
C	l	g
D	l	l

Your answer

11 Molten copper chloride is electrolysed.

What is made at the POSITIVE ELECTRODE (ANODE)? [1]

A Chlorine

B Copper

C Hydrogen

D Oxygen

Your answer

12 Barium carbonate contains the ions Ba^{2+} and CO_3^{2-} .

What is the formula for barium carbonate? [1]

A BaCO_3

B $\text{Ba}(\text{CO}_3)_2$

C Ba_2CO_3

D $\text{Ba}_2(\text{CO}_3)_2$

Your answer

13 Sodium carbonate reacts with dilute hydrochloric acid.

Which PRODUCTS are made? [1]

- A Sodium chloride + hydrogen + carbon dioxide**
- B Sodium chloride + water + carbon dioxide**
- C Sodium hydroxide + hydrogen + carbon dioxide**
- D Sodium hydroxide + water + carbon dioxide**

Your answer

14 How many covalent bonds can a carbon atom form? [1]

- A 2**
- B 4**
- C 6**
- D 8**

Your answer

15 What is the ionic equation for neutralisation? [1]



Your answer

BLANK PAGE

SECTION B

- 16 The diagram opposite shows a simplified version of the PERIODIC TABLE.**

The letters D, E, F, G, H and I are NOT the symbols for the elements shown.

- (a) Match each statement with one of the elements by writing D, E, F, G, H or I.**

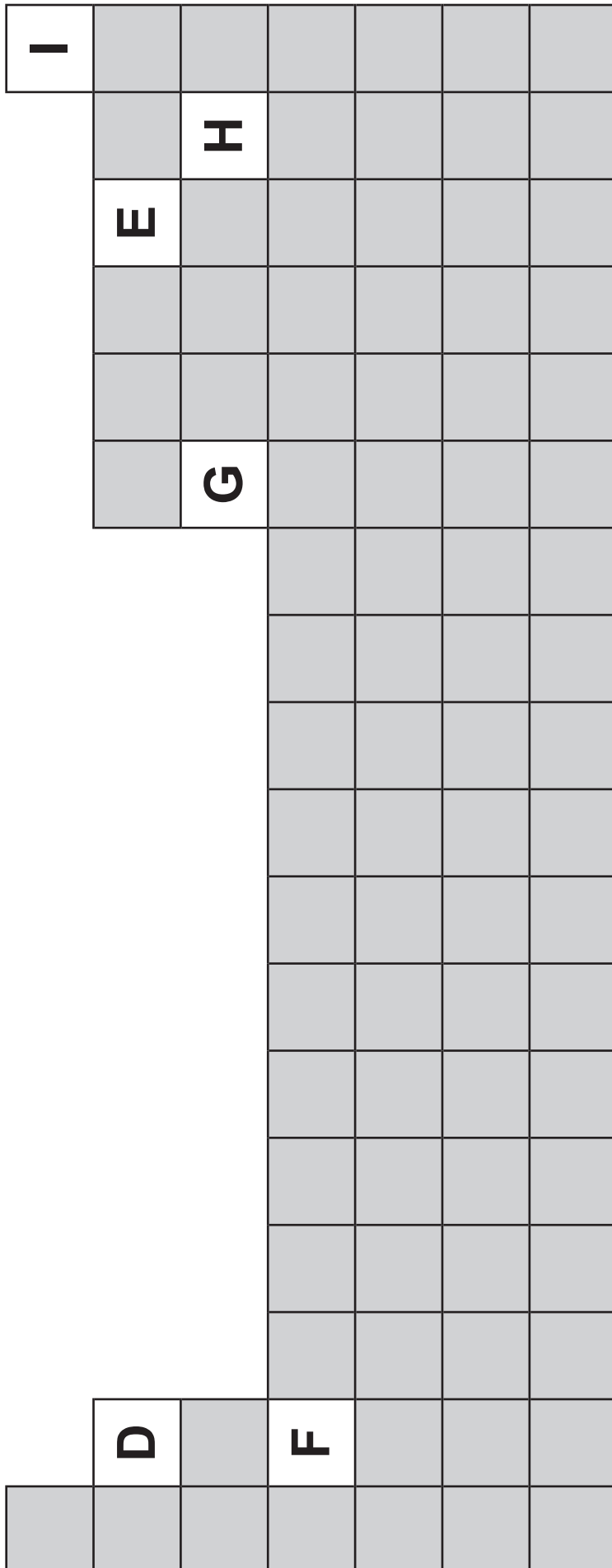
Use the Periodic Table on the data sheet.

- (i) An unreactive gas with a full outer shell of electrons.**

_____ **[1]**

- (ii) An element which is a solid at room temperature.**

_____ **[1]**



(iii) An element which forms an ion with a 1– charge.

_____ [1]

(iv) An element which is a gas at room temperature.

_____ [1]

(v) An element which has three shells of electrons.

_____ [1]

(b) Which TWO letters are elements in the same group?

_____ and _____ [1]

BLANK PAGE

17 A baker prepares some ingredients to make bread.

(a) The baker wants to understand how to use:

**an ACID (buttermilk)
and an ALKALI (sodium bicarbonate)**

to produce bubbles which cause the bread to rise.

Complete the sentences below.

Use words from the list.

**alkalis
carbon dioxide
gas
neutralisation
oxygen
precipitation
salts
solid**

Acids have a pH of below 7 and

_____ have a pH of above 7.

The acid and alkali react together

to form bubbles because a

_____ is produced.

The bubbles are _____ .

This is an example of a

_____ reaction. [4]

(b)

(i) Buttermilk is a liquid.

Describe how the baker can measure the pH of the buttermilk.

[2]

(ii) The table shows the pH of several acidic ingredients.

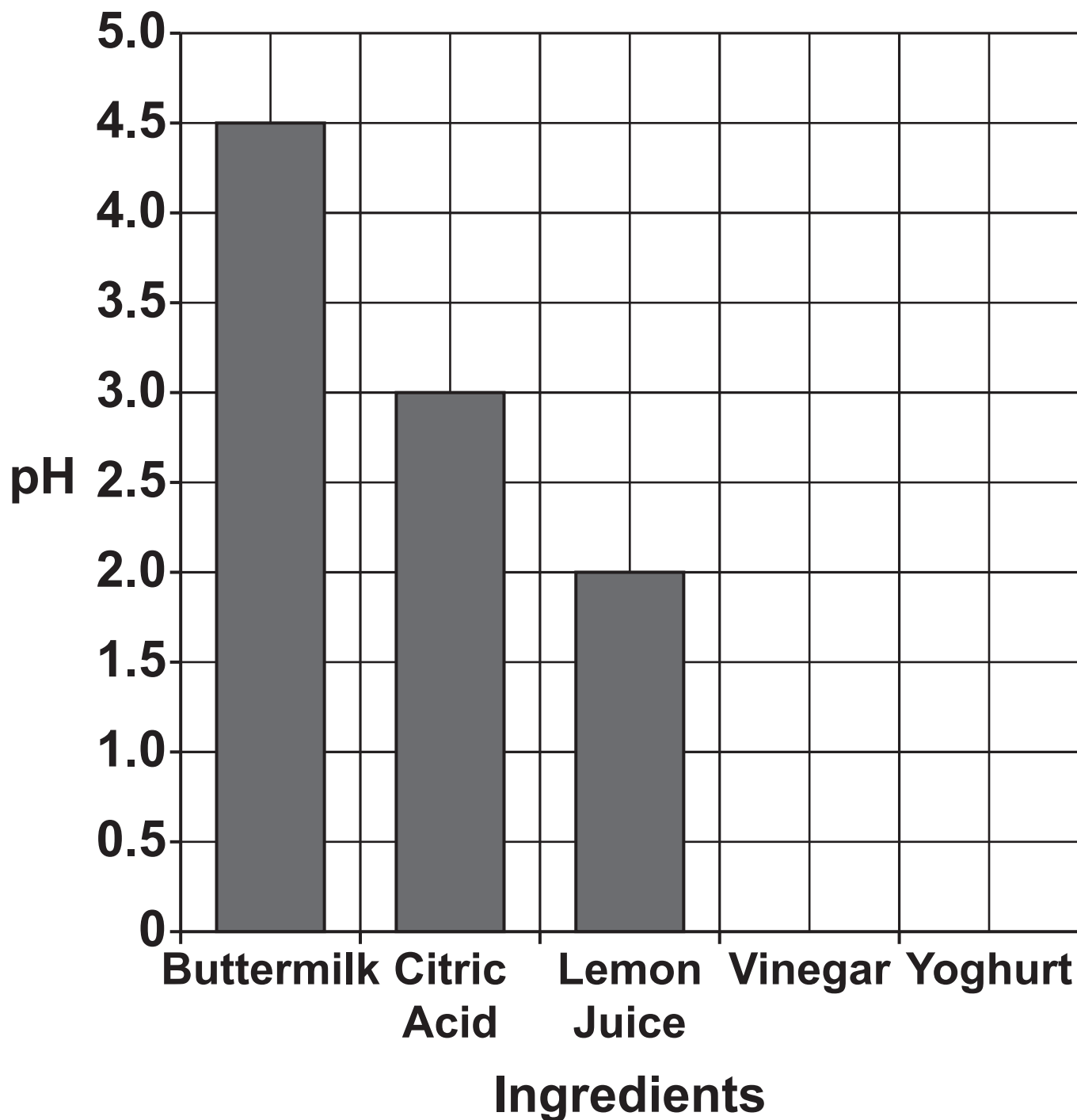
Ingredient	pH
Buttermilk	4.5
Citric acid	3.0
Lemon juice	2.0
Vinegar	2.5
Yoghurt	5.0

Which ingredient is the MOST acidic?

[1]

(iii) Use the data in the table to complete the bar chart.

Some of the bars have been done for you. [2]



(c) The formula of citric acid is $\text{C}_6\text{H}_8\text{O}_7$.

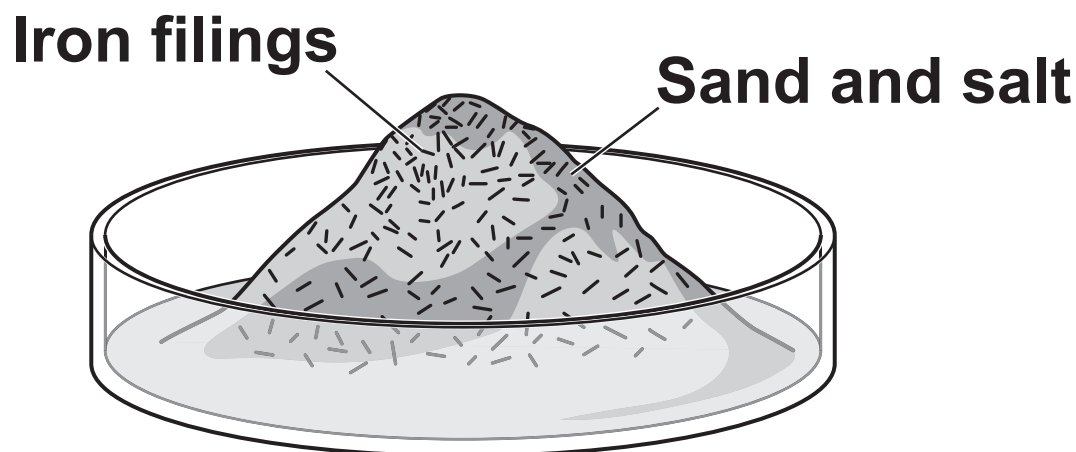
Calculate the relative formula mass of citric acid.

Relative atomic mass (A_r):

C = 12.0 H = 1.0 O = 16.0

**Relative formula
mass of citric acid = _____ [2]**

18* A student has a mixture of iron filings, sand and salt.



The student wants to collect pure samples of:

**iron filings
sand
salt.**

The table gives some information about iron filings, sand and salt.

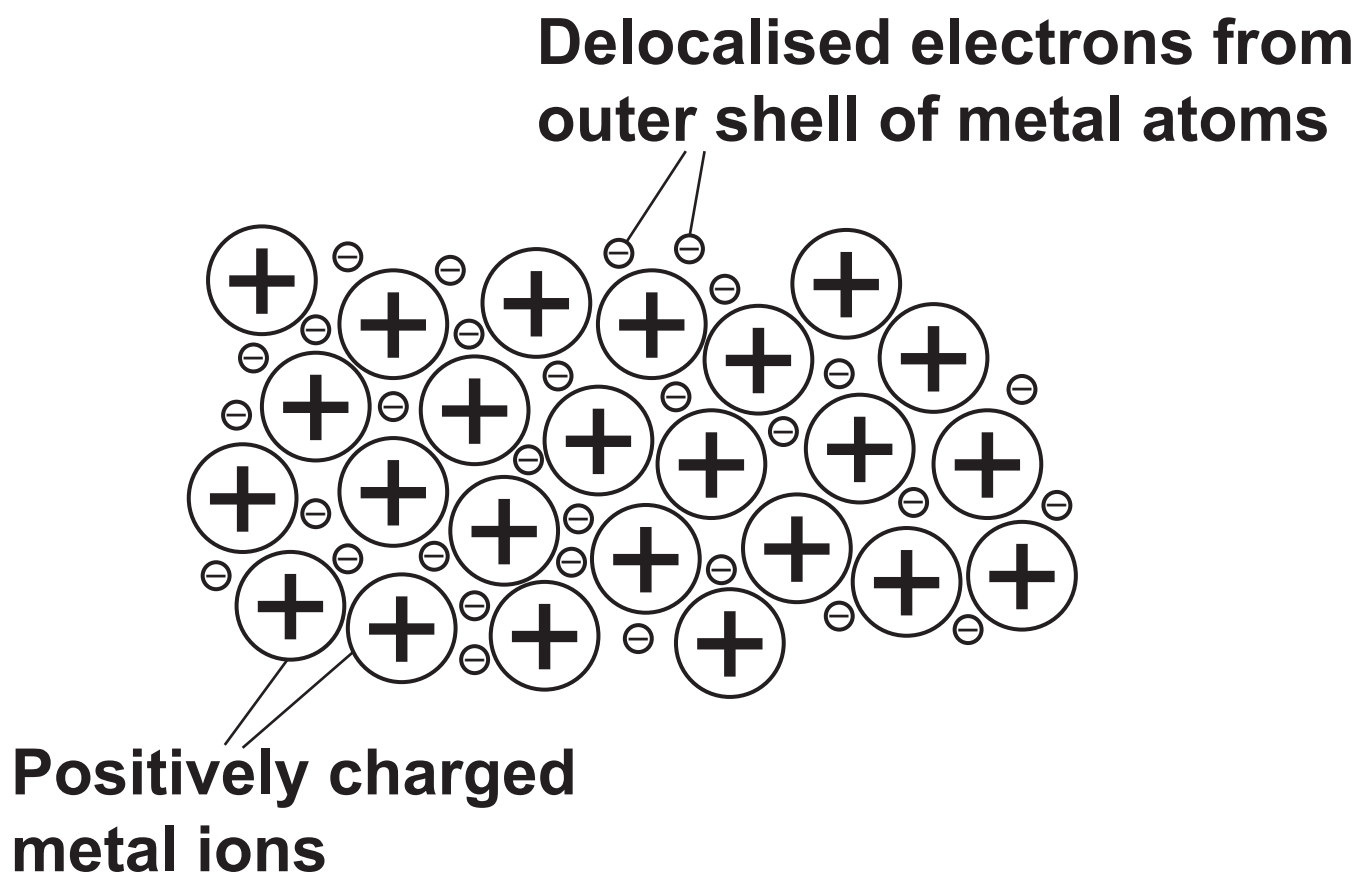
Substance	Soluble in water?	Magnetic?	Melting point (°C)
Iron filings	no	yes	1535
Salt	yes	no	801
Sand	no	no	1700

Describe a method the student could use to separate and collect pure iron filings, sand and salt.

You can include labelled diagrams to support your answer. [6]

Optional diagrams:

19 The diagram shows the structure of a metal.



(a) The table shows some information about three different metals.

Metal	Resistance to corrosion	Relative strength	Density
A	low	high	high
B	high	low	medium
C	very high	high	low

Which of these metals would you use for an aeroplane?

Explain your answer.

Use information from the table.

Metal _____

Explanation _____

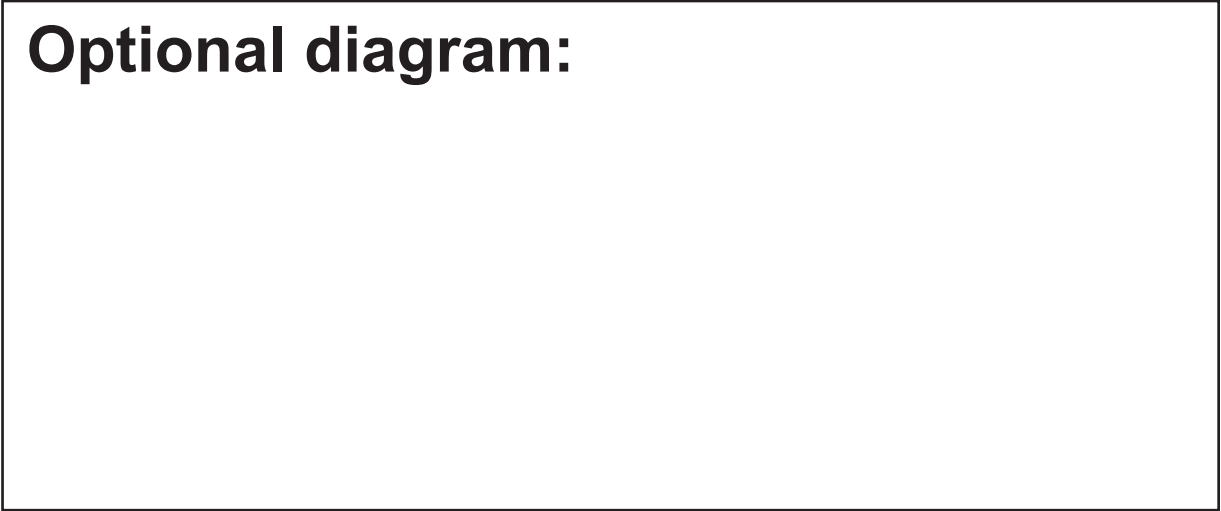
[3]

(b)

(i) Explain why alloys are harder than metals.

You can include a labelled diagram to support your answer.

Optional diagram:



[2]

(ii) Steel alloys can be used to make tools.

A sample of tool steel contains 0.17 g of carbon and 7.73 g of iron.

Calculate the PERCENTAGE OF CARBON in this alloy.

Give your answer to 3 significant figures.

Percentage of carbon = _____ % [4]

20 Atoms contain three subatomic particles: protons, neutrons and electrons.

(a) TABLE 20.1 shows some information about the subatomic particles.

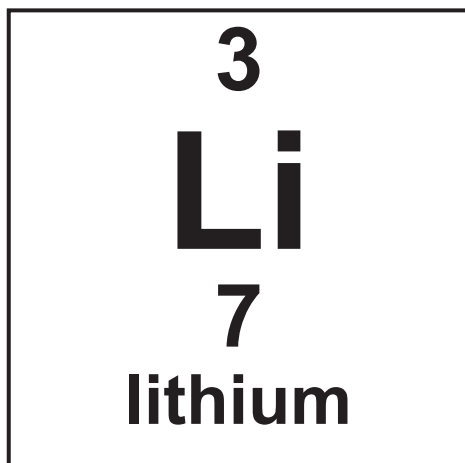
TABLE 20.1

Subatomic particle	Relative mass	Relative charge
Electron		
Neutron	1	
Proton		+1

Complete the table to show the relative mass and relative charge of each subatomic particle. [3]

(b) FIG. 20.1 shows some information about a lithium atom.

FIG. 20.1



Which statements about this lithium atom are TRUE, and which are FALSE?

Tick (✓) ONE box in each row. [2]

	TRUE	FALSE
The atom has 3 electrons		
The atom has 3 neutrons		
The atom has 4 protons		

- (c) Complete the sentence to explain how a lithium atom forms a lithium ion.**

Put a ring around each correct option.

The lithium atom gains / loses / shares one electron to form a negative / neutral / positive ion. [2]

(d)

- (i) Lithium can exist as different isotopes.**

Complete the sentence about isotopes.

Use words from the list.

**compound
electrons
element
ion
neutrons
protons**

Isotopes are atoms of the same

_____ with the same

number of _____ but a

different number of _____. [3]

(ii) TABLE 20.2 shows information about two different isotopes of lithium.

TABLE 20.2

ISOTOPE	MASS NUMBER
_____ Li _____	6
3 Li 7	_____

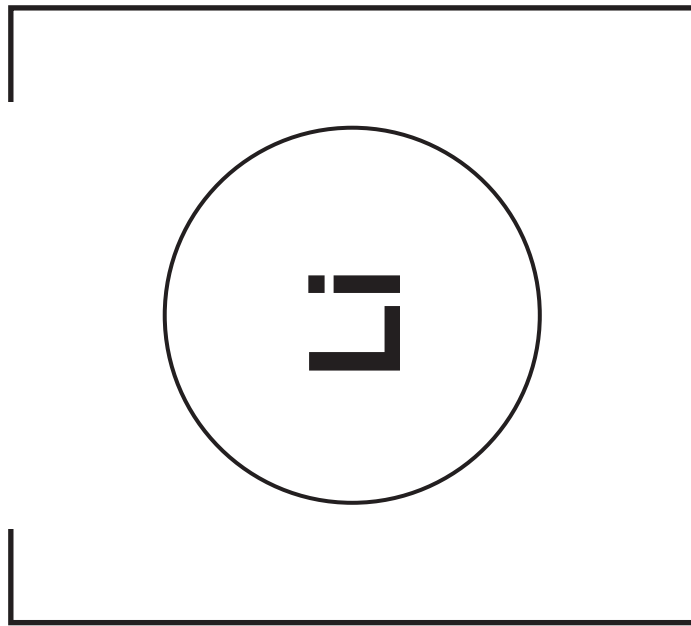
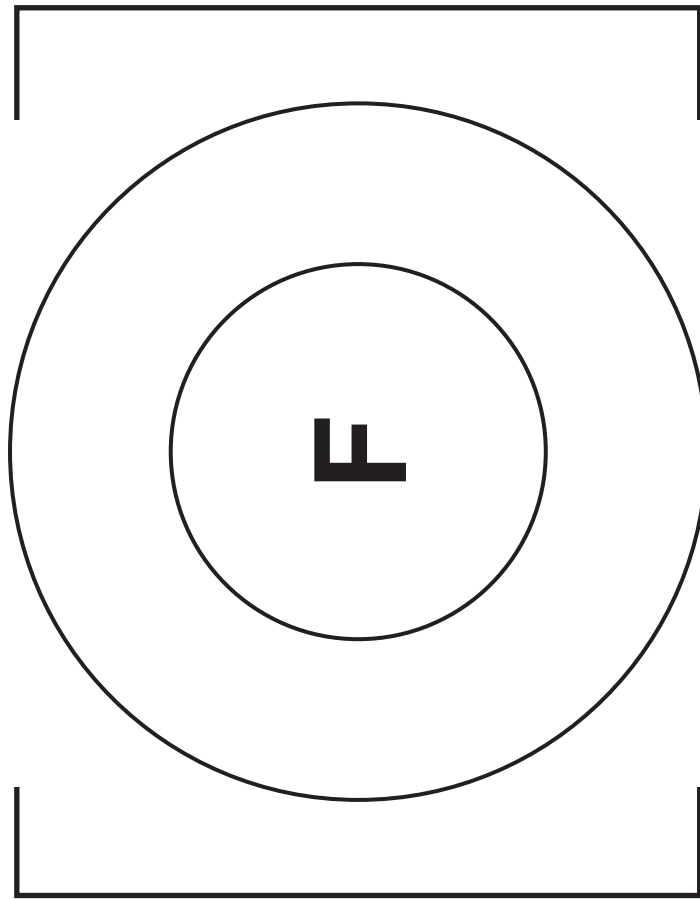
Complete the table. [2]

(e) Lithium can bond with fluorine to form lithium fluoride, LiF.

Complete the dot and cross diagram in FIG. 20.2 opposite to show the bonding in lithium fluoride.

Include the charges on the ions. [2]

FIG. 20.2

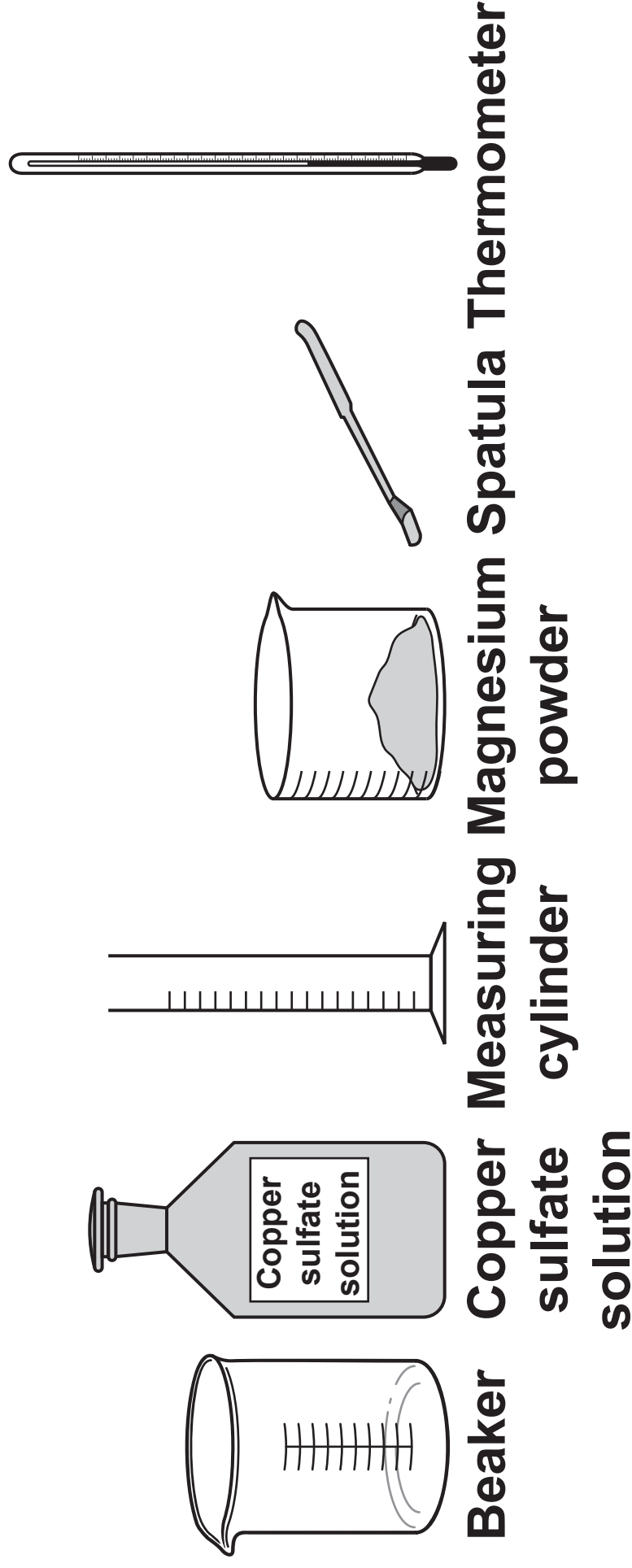


21 A student investigates the temperature change in a reaction.

(a)

(i) The student collects the apparatus needed to do their investigation, as shown in FIG. 21.1 opposite.

FIG. 21.1



Describe the experiment the student does to measure the temperature change.

You can include a labelled diagram to support your answer.

Optional diagram:

[4]

BLANK PAGE

- (ii) The student does the experiment three times and records their results in TABLE 21.1 opposite.

The student uses their results to determine if the reactions are endothermic or exothermic.

Tick (✓) ONE box in each row of TABLE 21.2 to show if each result is endothermic or exothermic. [1]

TABLE 21.2

Experiment	Endothermic	Exothermic
1		
2		
3		

TABLE 21.1

Experiment	Start temperature (°C)	End temperature (°C)	Temperature change (°C)
1	21	63	+42
2	21	7	−14
3	21	34	+13

(b) The student thinks they can improve their experiment to get more accurate results.

Suggest ONE improvement they can make to the apparatus shown in FIG. 21.1.

Explain why this will make their results more accurate.

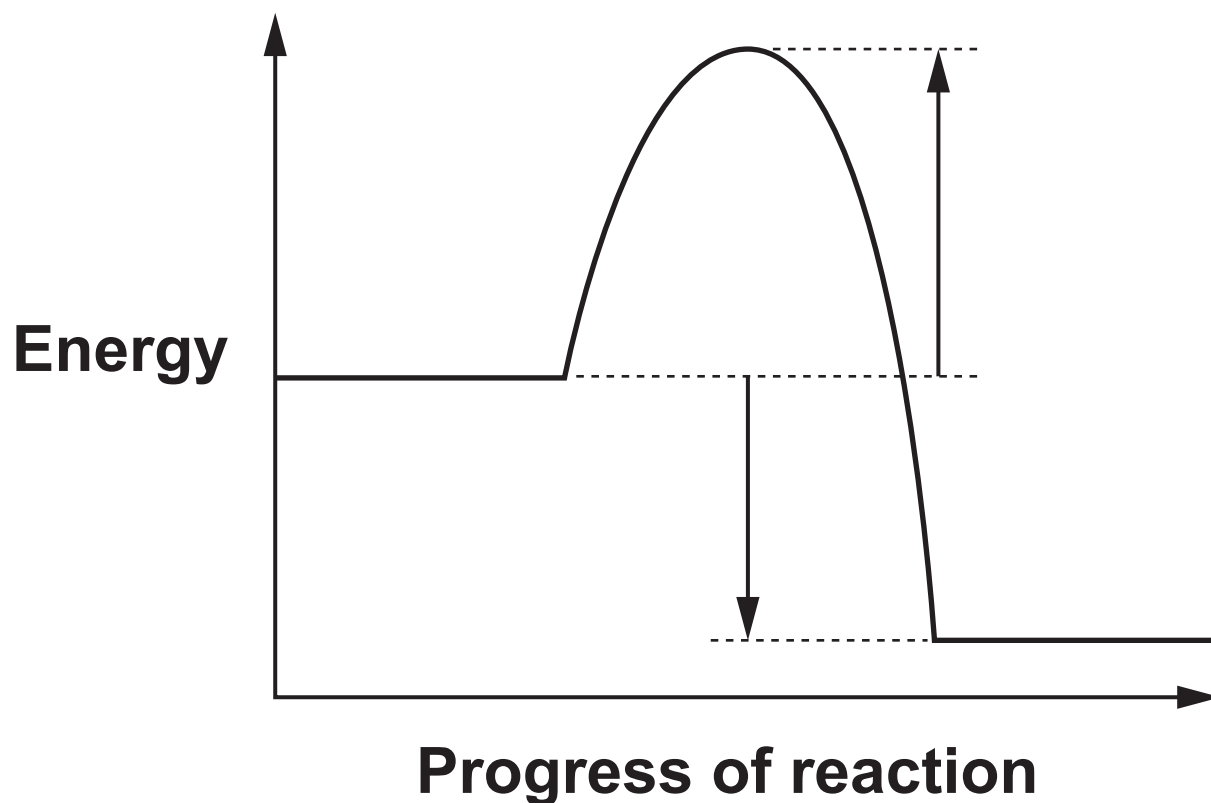
Improvement _____

Explanation _____

[2]

(c) FIG. 21.2 shows a reaction profile for one of the experiments.

FIG. 21.2



(i) Label on FIG. 21.2 the:

Reactants

Products

Activation energy

Energy change. [3]

(ii) Is the reaction shown in FIG. 21.2 endothermic or exothermic?

Tick (✓) ONE box.

Endothermic

☐

Exothermic

☐

Explain your answer.

[1]

22 Scientists are developing the use and application of nanoparticles.

(a) The table shows the diameters of four different particles.

PARTICLE	DIAMETER (nm)
A	97
B	2100
C	11
D	0.73

Which TWO particles are nanoparticles?

_____ and _____

Explain your answer.

[2]

(b)

(i) One use of nanoparticles is in medicines.

State ONE OTHER use of nanoparticles.

Explain why they can be used for that purpose.

Use _____

Explanation _____

[2]

- (ii) Some scientists are worried about the risks of using nanoparticles.**

Which statements are risks to humans of using nanoparticles?

Tick (✓) TWO boxes. [2]

They can block ultraviolet light.

☐

They can catalyse harmful reactions in the body.

☐

They can have antibacterial properties.

☐

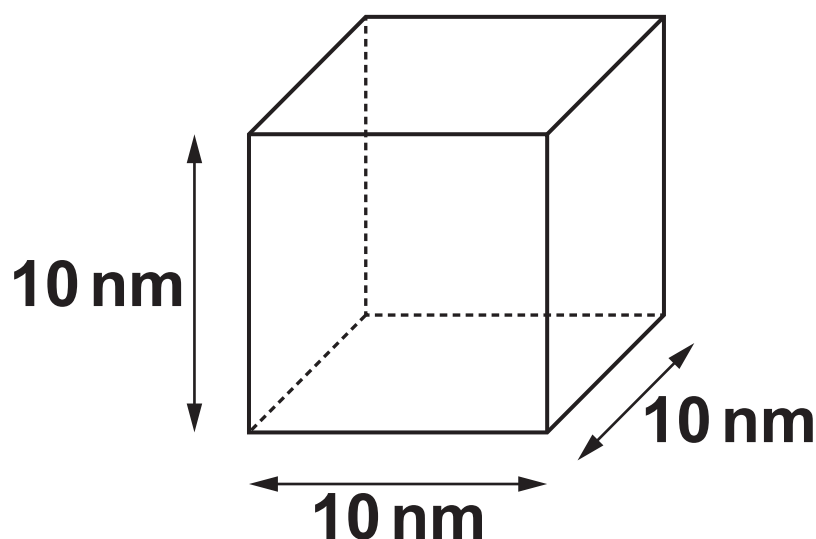
They can have toxic effects on the body.

☐

They have a high surface area to volume ratio.

☐

- (c) A cube-shaped nanoparticle has sides of length 10 nm. You may use a model to help you.



Calculate the surface area to volume ratio of the cube.

Use the equation:

$$\text{Surface area to volume ratio} = \frac{\text{surface area}}{\text{volume}}$$

Surface area to volume ratio = _____ [3]

BLANK PAGE

23 This question is about separation techniques.

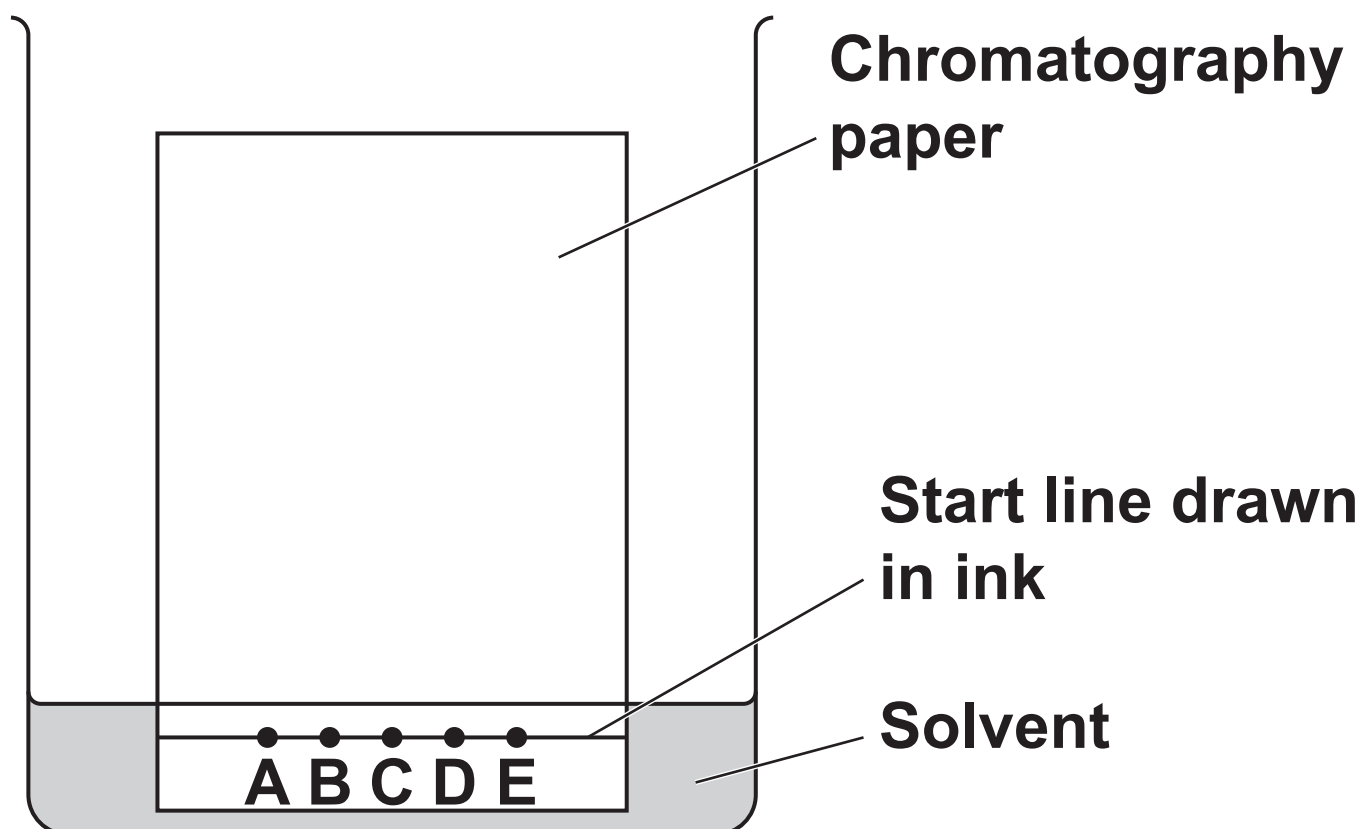
(a) What is meant by a PURE substance?

[1]

(b)

(i) A student does a chromatography experiment to separate some dyes.

The diagram shows how they set up their equipment.



The student's experiment will NOT work to separate the dyes.

Suggest TWO changes the student can make so that the experiment will work.

1 _____

2 _____

[2]

- (ii) The student calculates that a dye has an R_f value of 0.67.

The table shows the R_f value of different dyes.

Colour of dye	R_f value
Blue	0.15–0.20
Red	0.35–0.40
Pink	0.40–0.80
Yellow	0.75–0.80

Which dye did they use in their experiment?

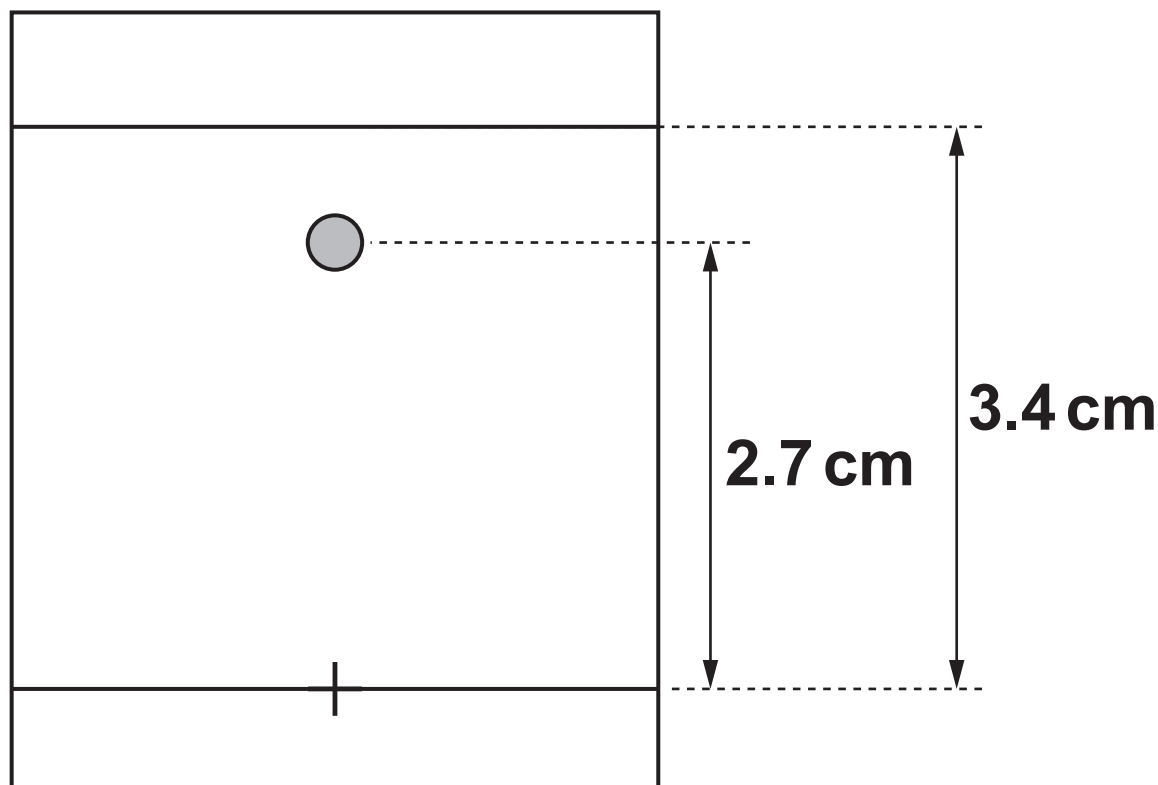
Explain your answer.

Dye _____

Explanation _____

[2]

(c) The diagram shows a chromatogram for another dye.



Calculate the R_f value of the dye.

Use the equation:

$$R_f = \frac{\text{distance travelled by the substance}}{\text{distance travelled by the solvent}}$$

R_f value = _____ [2]

(d) A student separates substance S from a mixture of substances S and T.

The starting mass of the mixture is 9.7 g and they obtain 5.1 g of pure substance S.

Calculate the mass of pure substance S that could be separated from 15.0 g of the mixture.

Mass of pure substance S = _____ g [2]

END OF QUESTION PAPER

EXTRA ANSWER SPACE

If you need extra space use these lined pages. You must write the question numbers clearly in the margin.



Copyright Information

OCR is committed to seeking permission to reproduce all third-party content that it uses in its assessment materials. OCR has attempted to identify and contact all copyright holders whose work is used in this paper. To avoid the issue of disclosure of answer-related information to candidates, all copyright acknowledgements are reproduced in the OCR Copyright Acknowledgements Booklet. This is produced for each series of examinations and is freely available to download from our public website (www.ocr.org.uk) after the live examination series.

If OCR has unwittingly failed to correctly acknowledge or clear any third-party content in this assessment material, OCR will be happy to correct its mistake at the earliest possible opportunity.

For queries or further information please contact The OCR Copyright Team, The Triangle Building, Shaftesbury Road, Cambridge CB2 8EA.

OCR is part of Cambridge University Press & Assessment, which is itself a department of the University of Cambridge.